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A polyatomic ion, therefore, is an ion that is made of more than one atom. An example of a polyatomic ion is the sulfate ion,  $\text{SO}_4^{2-}$ . Sulfate is composed of one sulfur atom and four oxygen atoms and overall sulfate has a negative two charge.

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7. Write formulas for the following ionic compounds. Note that each includes a polyatomic ion. a) lithium phosphate b) ammonium oxide c) barium hydroxide  $\text{Li}_3\text{PO}_4$   $(\text{NH}_4)_2\text{O}$   $\text{Ba}(\text{OH})_2$



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2 d) calcium cyanide e) sodium chlorate f) potassium sulfate  
Ca(CN)<sub>2</sub> NaClO<sub>3</sub> K<sub>2</sub>SO<sub>4</sub>. In question 3, you were asked the  
difference between calcium nitride and ...

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dot diagrams will continue to be useful throughout the unit, and the semester. as students are introduced to covalent bonding and continue to recognize periodicity in whether atoms gain, lose or share electrons to form ...

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Naming Ions To write an ion you write the symbol of the atom

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and put the charge in the upper right. ... Ionic bonding is when two ions of opposite charge attract and bond to each other forming an ionic compound.